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Properties

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Chemistry

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Properties
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Molality and

Colligative

Properties

~~Solute, Solvent,~~

~~& Solution~~

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~~Chemistry~~

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Properties

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Review Study

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*Types of
Solution | What
is a solution?*

Chemistry

Solution Solvent

Solute -

Definition and

Difference

13.1

Introduction to

Colligative

Properties, the

van't Hoff

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factor, and
Molality ~~Aqueous~~
~~Solution~~
Chemistry

Colligative

Properties

calculate all of
them! Worked out
problem(s).

Chapter 14 -

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**Kinetics: Part 1
of 17**

Chapter 13 -

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~~Colligative~~

~~Properties~~

~~Explained~~

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Magnet Brains

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vocabulary,

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Chapter 16:

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Properties of
Solutions.

Solution

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compositions of
the solvent and
the solute
determine

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whether a substance will dissolve. The factors that determine how fast a substance dissolves are stirring (agitation) temperature the surface area of the dissolving particles 16.1

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Questions ...

formation of

solution can be

either

exothermic or

endothermic;

exothermic

processes are

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spontaneous;
solution will
not form if
enthalpy is too
endothermic; H_3
has to be
comparable to $H_1 + H_2$. Ionic
substances
cannot dissolve
in nonpolar
liquids; Polar
liquids do not
form solutions

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with nonpolar
liquids; 13.1.2

Solution

Formation,

Spontaneity, and

Disorder

*13.S: Properties
of Solutions*

(Summary) -

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Characteristics

Types

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Properties. What is a Solution? A solution is a homogeneous mixture of two or more components in which the particle size is smaller than 1 nm. Common examples of solutions are the sugar in

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water and salt
in water
solutions, soda
water, etc. In a
solution, all
the components
appear as a
single phase.

*Solution -
Definition,
Properties,
Types, Videos &
Examples*

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Properties of
Solutions
Chemistry, The
Central Science
, 10th edition
Theodore L.

Brown; H. Eugene
LeMay, Jr.; and
Bruce E.

Bursten.

Solutions

Solutions •

Solutions are

homogeneous

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mixtures of two or more pure substances. • In a solution, the solute is dispersed uniformly throughout the solvent .

Chapter 13

Properties of

Solutions

Small-Scale

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Laboratory

Manual, Lab 26 •

Laboratory

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16.2, Assessment

16.1 • Go

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Connecting to

Your World

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Solutions 471

16.1 ...

16.1 Properties

of Solutions 16

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In a solution in which carbon dioxide is dissolved in water, the water is the solvent and the carbon dioxide is the solute. Two important concepts in studying chemical solutions are

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solution
concentration
and solubility
equilibrium.

Properties of solutions as a whole are called colligative properties. How to recognize different types of solutions.

Chemical

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*Solutions (with
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Chemistry. ...

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relationships of
basic chemical
principles and
theories to
properties of
substances,
their reactivity

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and uses. **Properties Of**

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the quality of
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A.P. Chemistry

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Ch. 11,
Solutions

Name _____

MULTIPLE CHOICE.

Choose the one alternative that best completes the statement or answers the question. 1)

Formation of solutions where the process is

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endothermic can
be spontaneous
provided that
 . A) the

solvent is a gas
and the solute
is a solid

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